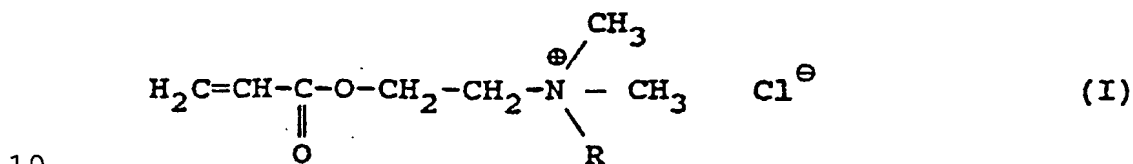


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PROCESS AND MANUFACTURE OF AQUEOUS SOLUTIONS OF
UNSATURATED QUATERNARY AMMONIUM SALTS

The present invention relates to the
5 manufacture of aqueous solutions of unsaturated
quaternary ammonium salts (hereinafter denoted
quaternary salts) corresponding to the following
formula (I):



in which R represents methyl or benzyl,
by reaction, in the presence of water, of N,N-dimethyl-
aminoethyl acrylate (DAMEA) with a quaternizing agent
15 of formula (II):



in which R is as defined above.

20 Aqueous solutions of quaternary salts (I) are
used to prepare polymers intended to act as cationic
flocculents in water treatment.

European patent EP-B-250 325 discloses a
process for the preparation of aqueous solutions of
25 quaternary salts, including those of formula (I)
according to which process, in the presence of at least
one polymerization inhibitor:

- in a first stage (a), DAMEA is reacted with 5 to
20% by weight of the amount by weight of the
30 quaternizing agent necessary for the reaction or,
according to an alternative form (a'), with 5 to
20% by weight, with respect to the weight of the
DAMEA, of an aqueous solution of quaternary salts,
which solution comprises from 50 to 85% by weight
35 of quaternary salts; and

- in a second stage (b), water and the quaternizing agent are continuously added until the desired concentration of quaternary salts in the water is obtained.

5 During stages (a) and (b), the temperature is maintained at a value of between 30 and 60°C. Furthermore, during stages (a) and (b) and in particular near the end of the reaction, a stream of oxygenated gas is maintained in the reaction medium
10 such that the ratio by volume (or volumetric throughput) of total gas at the outlet of the reactor to the volume (or volumetric throughput) of oxygen introduced at the inlet of this same reactor is less than 100.

15 This process makes it possible to prepare aqueous solutions of quaternary salts which have a stability at ambient temperature of greater than one year. However, a particularly high content of impurities, in particular of $\text{CH}_2=\text{CH}-\underset{\text{O}}{\underset{\parallel}{\text{C}}}-\text{O}-\text{R}$, of

20 $\text{CH}_2=\text{CH}-\underset{\text{O}}{\underset{\parallel}{\text{C}}}-\text{OH}$ and of DAMEA, is found in these solutions.

In addition, this process requires relatively long reaction times, which represents an obvious economic disadvantage.

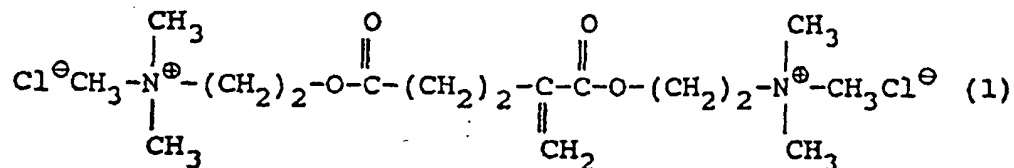
A process intended to reduce the formation of
25 the impurities during the quaternization reaction was then provided in international application WO 89/07 588. In accordance with this process, the reaction is carried out at a temperature of between 10 and 80°C, and

- 30 (a) in a first stage, all or a portion of the quaternizing agent necessary for the reaction is introduced into the reactor, this agent being in the liquid state under the reaction conditions,
(b) subsequently, the DAMEA is added, and
35 (c) as soon as 0 to 30% of the stoichiometry of the DAMEA has been introduced into the reactor, the

remainder of the quaternizing agent, the remainder of the DAMEA and the water are continuously and simultaneously added until the desired concentration of quaternary salts is obtained,

5 (d) and, in the case where the quaternizing agent is introduced in the gaseous state at the reaction temperature, the reaction is carried out in the presence of oxygen and a pressure is applied so that the quaternizing agent is liquid at the
10 reaction temperature, and, at the end of the reaction, the pressure is gradually reduced to atmospheric pressure and simultaneously a ratio as volumetric throughput of total gas at the outlet of the reactor to the volumetric throughput of
15 oxygen introduced into the reactor of less than 100 is imposed.

The above process according to WO 89/07 588 introduces significant improvement to the process according to EP-B-250 325. However, it transpired that
20 the purity with which the quaternary salts are obtained is still insufficient. Thus, during the reaction of DAMEA with CH₃Cl in aqueous medium, resulting in the salt also denoted subsequently by the abbreviation ADAMQUAT MC, the dimer of ADAMQUAT MC, represented by
25 the formula (1):



is formed as impurities, in addition to acrylic acid
30 (AA), formed by hydrolysis of DAMEA.

By virtue of a series of tests of reactivity with regard to polymerization, it was possible to demonstrate that these impurities affected the quality of the cationic polymers derived from ADAMQUAT.

35 The applicant company has thus looked for operating conditions for the preparation of aqueous

solutions of the salt of formula (I) which are capable of minimizing the abovementioned impurities, so as to provide a salt (I) in aqueous solution of very high analytical quality.

5 This novel process, which thus forms the subject matter of the present invention, is characterized in that:

10 (a) the reaction is carried out in a closed reactor, which comprises 5-60% of the amount by weight of DAMEA necessary for the reaction and which has been pressurized by air or depleted air to 0.5 to 3 bar, by continuously introducing, at a temperature of 35 to 65°C, in particular of 40 to 60°C, on the one hand, the quaternizing agent (II) and, on the other hand, the water and finally the remaining DAMEA, until the desired concentration of salt (I) in the water is obtained,

20 • the start of the introduction of the water beginning when 0-30%, in particular 10-20%, of the amount by weight of the quaternizing agent (II) necessary for the reaction has been added;

25 • the start of the introduction of the remaining DAMEA beginning when 20-80%, in particular 30-70%, of the amount by weight of the quaternizing agent (II) necessary for the reaction has been added; and

30 • it being possible for the pressure at the end of the reaction to reach 9 bar, in particular 4 to 7 bar; then

35 (e) the reactor is depressurized while keeping the oxygen content constant by simultaneous introduction of air and, after returning to atmospheric pressure, the residual quaternizing agent is removed, for example by stripping with air.

 In accordance with other specific characteristics of the process according to the invention:

- the quaternizing agent is introduced over a period of time of 1-7 hours, the water over a period of time of 1-8 hours and the remaining DAMEA over a period of time of 2-8 hours;
- 5 - the reaction is carried out with a molar ratio of the quaternizing agent to the DAMEA of 1 to 1.1, preferably of 1 to 1.05;
- the reaction is carried out with a mean ratio of water/quaternizing agent throughput of 0.2-1.5, in particular of 0.4-1, a mean ratio of remaining DAMEA/quaternizing agent throughput of 2.5-5, in 10 particular 3-4, and a mean ratio of water/remaining DAMEA throughput of 0.2-1.2, in particular of 0.3-0.9.

15 The process according to the invention makes it possible in particular to prepare aqueous solutions having a concentration of quaternary salts (I) of 50 to 85% by weight and comprising very low amounts of impurities, as illustrated in Table 1 below.

20 Furthermore, the process according to the present invention can be carried out in the presence of at least one stabilizer, which can be chosen from 3,5-di(tert-butyl)-4-hydroxytoluene, hydroquinone methyl ether, hydroquinone, catechol, tert-butylcatechol, phenothiazine and mixtures of these 25 stabilizers, the content of stabilizing agent(s) being in particular from 20 to 2 000 ppm, preferably from 100 to 1 200 ppm, with respect to the aqueous solution of quaternary salt (I).

30 In addition, at least one sequestering agent for metals chosen in particular from diethylenetriaminepentaacetic acid, the pentasodium salt of diethylenetriaminepentaacetic acid, N-(hydroxyethyl)-ethylenediaminetriacetic acid and the trisodium salt of 35 N-(hydroxyethyl)ethylenediaminetriacetic acid can be added to the reaction medium, the content of sequestering agent(s) being in particular from 1 to 100 ppm, preferably from 5 to 30 ppm, with respect to the aqueous solution of quaternary salt (I).

Generally, the sequestering agents are added in the form of an aqueous solution as they are generally available in this form. Thus, the pentasodium salt of diethylenetriaminepentaacetic acid sold under the name
5 Versenex 80 is provided in the form of an approximately 40% by weight aqueous solution.

The following examples illustrate the present invention without, however, limiting the scope thereof. From these examples, the percentages are by weight,
10 unless otherwise indicated.

EXAMPLE 1:

200 g of DAMEA (i.e. 46.6% of all the DAMEA) were charged to a 1 l jacketed glass reactor, specially
15 designed to withstand pressure, equipped with a temperature probe, with a gas/liquid specific stirrer (turbine with a hollow shaft), with a valve tared at 10 bar, with a bursting disc and with dip pipes for the introduction of the various reactants. The reactor was
20 closed and then pressurized with 1 bar of depleted air. Stirring and heating were begun.

As soon as the temperature reached 40°C (process temperature = 47°C), the introduction of CH₃Cl was begun at a throughput of 70 g/h. When 35 g of CH₃Cl
25 were introduced, the introduction of water was begun at a throughput of 28.6 g/h. After reacting for 1 h, the CH₃Cl throughput was brought back to 20.9 g/h. The introduction of the remainder of the DAMEA (i.e. 229 g) was begun at a throughput of 76.3 g/h after reacting
30 for 1.5 h. At the end of the reaction, the reactor was brought back to atmospheric pressure using the following protocol:

- degassing the excess CH₃Cl for 30 minutes with simultaneous introduction of air into the charge
35 (throughput: 3 Sl/h);
- gradual return to atmospheric pressure, and the traces of CH₃Cl were removed by stripping with air (throughput: 5 Sl/h) for 30 minutes.

The reactor was subsequently cooled and then emptied. 710 g of ADAMQUAT MC 80 were recovered and were analyzed by high performance liquid chromatography (HPLC) to determine the contents of AA and of compound
5 (1). The results are reported in Table 1.

The durations of the various phases of the reaction were as follows:

- CH₃Cl introduction : 5.25 h
 - H₂O introduction : 5 h
 - 10 - DAMEA introduction : 3 h
 - degassing : 0.5 h
 - stripping : 0.5 h
- i.e. a total duration of approximately 6.75 h.

The throughput ratios used were:

- 15 - H₂O/CH₃Cl : 0.41 for the first hour, then 1.37
for the remainder of the reaction;
- H₂O/DAMEA : 0.37
- DAMEA/CH₃Cl : 3.64.

20 EXAMPLE 2:

The preparation was carried out as in Example 1, except that the CH₃Cl throughput was increased.

The results are also reported in Table 1.

Table 1

[CH₃Cl]/[DAMEA] = 1.05; T = 47°C; Maximum pressure: 6 bar; DAMEA: 429 g (3 mol); CH₃Cl = 15 g (3.15 mol)

Example	Mass of the crude reaction mixture (g)	Duration of introduction of CH ₃ Cl (h)	H ₂ O/ DAMEA	DAMEA/ CH ₃ Cl	H ₂ O/ CH ₃ Cl	HPLC analysis (ppm)		
						AA	ADAMQUAT MC dimer of formula (1)*	QUATS (%)
1	710	5.25	0.37	3.65	0.4 then 1.37	426	407	81
2	724	4	"	"	"	418	277	82.8

* Content of ADAMQUAT MC dimer of formula (1), expressed arbitrarily as AA